

Name \_\_\_\_\_

## Ch 5 CP Study Guide

Who was the first person to develop the Periodic table?

How did he arrange the elements on the periodic table?

Define Periodicity

Explain why there were some question marks in the place of elements on the early periodic table?

How did Moseley rearrange the periodic table in 1911?

Why didn't it make sense that Te had to come before I on the early periodic table?

The density of Ca is  $1.55 \text{ g/cm}^3$  and the density of Ba is  $3.51 \text{ g/cm}^3$ , estimate the density of Sr. What is your percent error, when the actual density is  $2.66 \text{ g/cm}^3$ ?

Why were the noble gases left off early periodic tables?

What are the groups and periods on the periodic table?

What element is in:

Period 3, Group 15 \_\_\_\_ How many valence electrons does it have? \_\_\_\_

Period 2 Group 13 \_\_\_\_ How many valence electrons does it have? \_\_\_\_

Period 4, Group 2 \_\_\_\_ How many valence electrons does it have? \_\_\_\_

What is atomic radius and how is it measured?

Explain what happens to atomic radius down a group on the periodic table? Across the period?  
EXPLAIN WHY IN BOTH CASES

What is ionization energy?

Explain what happens to ionization energy down a group on the periodic table? Across the period?  
EXPLAIN WHY IN BOTH CASES

What is electronegativity?

Explain what happens to electronegativity down a group on the periodic table? Across the period?  
EXPLAIN WHY IN BOTH CASES

What is the most electronegative element?

Define cation and anion.

How does the size of a cation compare to its neutral atom?

How does the size of an anion compare to its neutral atom?

List the properties AND location on the periodic table of the following types of elements

Alkali Metals   Alkaline Earth Metals   Transition Metals   Halogens   Noble Gases   Lanthanides/Actinides